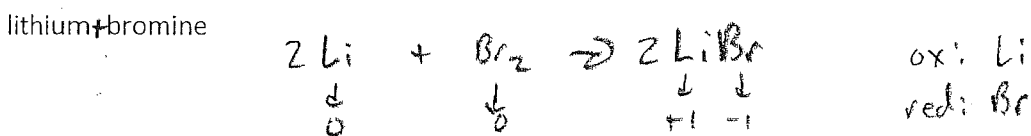
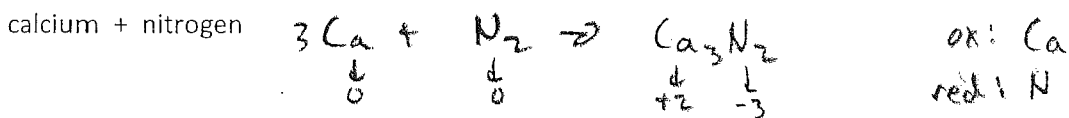
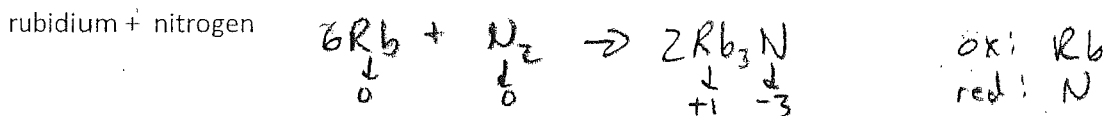
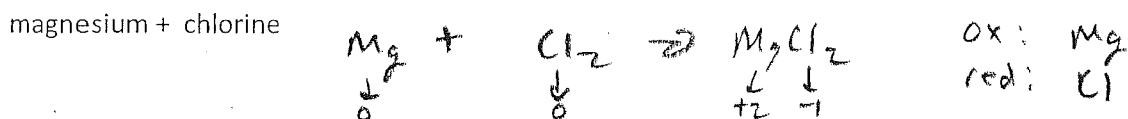
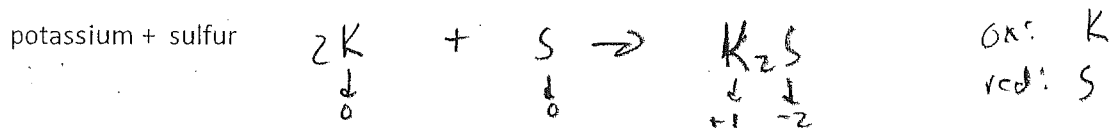
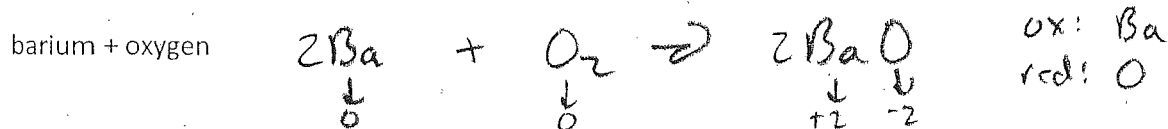
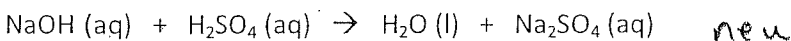
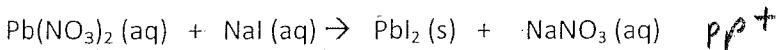
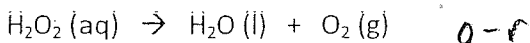
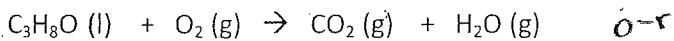
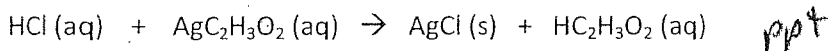


Oxidation-Reduction Reactions/Types

H) Write a balanced oxidation-reduction equation for each of the metals combining with a nonmetal. Identify what got oxidized and what got reduced.



He) Classify the following reactions as either: precipitation, neutralization, or oxidation-reduction.



Neutralization Reactions

Complete the word equations, then write balanced formula equations with physical states for the following neutralization reactions. Assume all reactants are aqueous.

(16) hydrochloric acid + rubidium hydroxide → water + rubidium chloride



perchloric acid + sodium hydroxide → water + sodium perchlorate



calcium hydroxide + hydrobromic acid → water + calcium bromide



sulfuric acid + cesium hydroxide → water + cesium sulfate



(8) N) Indicate the reactants needed to make the following salts and water. Id the acid and the base.

